

**Background:** Sadly last week was not the last week with stoichiometry. We have another week set aside to fully master this concept. Just think at how good you are going to be at doing mass balance problems. You can probably feel yourself getting cooler by the minute.

We will re-explore this concept by focusing on the mole to mole ratio that is inherent in the problem. It is at the heart of the balanced chemical equation. We need to remember that the balanced chemical equation tells us the relative number of atoms or molecules in a reaction. This does not represent the mass of atoms or molecules in a reaction.

We will continue to practice with stoichiometry. For such a fundamental concept, there are a lot of ways to phrase a question to make look like an entirely new concept. By the time we are done with this, you should be able to see all of these questions as basically the same.

**Objectives:**

1. Determine mass relationships between species in a balanced chemical reaction.
2. Identify the limiting reagent in a chemical reaction.
3. Discern between theoretical and actual yield of a chemical reaction.
4. Calculate the percent yield of a chemical reaction.

**Reading:** Zumdahl Chapter 9